



basic education

Department:
Basic Education
REPUBLIC OF SOUTH AFRICA

NATIONAL SENIOR CERTIFICATE

GRADE 10

PHYSICAL SCIENCES: CHEMISTRY (P2)

NOVEMBER 2019

MARKS: 150

TIME: 2 hours

This question paper consists of 13 pages and 2 data sheets.



INSTRUCTIONS AND INFORMATION

1. Write your name and class (e.g. 10A) in the appropriate spaces on the ANSWER BOOK.
2. This question paper consists of EIGHT questions. Answer ALL the questions in the ANSWER BOOK.
3. Start EACH question on a NEW page in the ANSWER BOOK.
4. Number the answers correctly according to the numbering system used in this question paper.
5. Leave ONE line between subquestions, e.g. between QUESTION 2.1 and QUESTION 2.2.
6. You may use a non-programmable calculator.
7. You may use appropriate mathematical instruments.
8. You are advised to use the attached DATA SHEETS.
9. Show ALL formulae and substitutions in ALL calculations.
10. Round off your FINAL numerical answers to a minimum of TWO decimal places.
11. Give brief motivations, discussions, etc. where required.
12. Write neatly and legibly.



QUESTION 1: MULTIPLE-CHOICE QUESTIONS

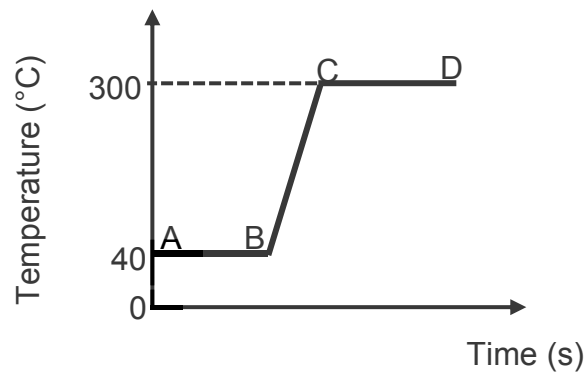
Various options are provided as possible answers to the following questions. Choose the answer and write only the letter (A–D) next to the question numbers (1.1 to 1.10) in the ANSWER BOOK, e.g. 1.11 E. Each question has only ONE correct answer.

1.1 Nitrogen gas is an example of a/an ...

- A element.
- B compound.
- C heterogeneous mixture.
- D homogeneous mixture.

(2)

1.2 The heating curve, not drawn to scale, of a compound is shown below.



During which section(s) on the curve will the potential energy of the compound INCREASE?

- A BC only
- B CD only
- C AB and CD
- D AB, BC and CD

(2)

1.3 Elements in the periodic table are arranged in order of increasing ...

- A mass number.
- B number of protons.
- C number of neutrons.
- D number of nucleons.

(2)

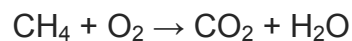


1.4 Avogadro's number is equal to the number of ...

- A atoms in 1 mole CO.
- B atoms in 1 mole Br₂.
- C molecules in 1 mole Au.
- D molecules in 1 mole N₂.

(2)

1.5 The unbalanced equation for a chemical reaction is shown below.



Which ONE of the following represents the coefficients of reactants and products in the BALANCED equation?

	CH ₄	O ₂	CO ₂	H ₂ O
A	2	1	2	1
B	1	1	2	2
C	1	2	1	2
D	2	2	1	1

(2)

1.6 A covalent bond forms ...

- A between metal and non-metal atoms.
- B through electron transfer.
- C through sharing of electrons.
- D between positive and negative ions.

(2)

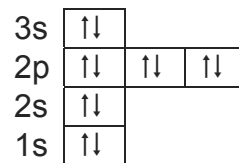
1.7 The reaction between hydrogen chloride (HCl) and sodium hydroxide (NaOH) is an example of a/an ... reaction.

- A redox
- B acid-base
- C precipitation
- D gas forming

(2)



1.8 Consider the Aufbau diagram of an element below.



The element is a/an ...

- A halogen.
- B noble gas.
- C alkali metal.
- D alkaline-earth metal. (2)

1.9 Which ONE of the following equations represents a precipitation reaction?

- A $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- B $\text{NaCl} + \text{HNO}_3 \rightarrow \text{NaNO}_3 + \text{HCl}$
- C $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$
- D $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$ (2)

1.10 The air surrounding the Earth is the ...

- A biosphere.
 - B lithosphere.
 - C atmosphere.
 - D hydrosphere. (2)
- [20]



QUESTION 2 (Start on a new page.)

Most elements found on Earth are metals. All metals have a metallic structure.

- 2.1 For a metallic structure, write down the:
- 2.1.1 Type of particles present (2)
- 2.1.2 Type of chemical bond formed between particles (1)
- 2.2 In which region on the periodic table are metals found? Choose from LEFT or RIGHT. (1)
- 2.3 Write down the SYMBOL of a metal which is a liquid at room temperature. (1)
- 2.4 To prevent iron from rusting it is often mixed with other metals, e.g. stainless steel is a mixture of iron, chromium and nickel.
- 2.4.1 Is stainless steel a HOMOGENEOUS or HETEROGENEOUS mixture? Give a reason for the answer. (2)
- 2.4.2 Rusting is the formation of iron(III) oxide. Write down the FORMULA of iron(III) oxide. (1)
- 2.5 The physical properties of two substances, **A** and **B**, are shown in the table below.

SUBSTANCES	DENSITY (g·cm ⁻³)	ELECTRICAL CONDUCTIVITY	THERMAL CONDUCTIVITY
A	4,94	Poor	Poor
B	7,87	Good	Good

- 2.5.1 Which ONE of the substances, **A** or **B**, has the smaller mass to volume ratio? Give a reason for the answer. (2)
- 2.5.2 Explain the difference between *electrical conductivity* and *thermal conductivity*. (2)
- 2.5.3 Which ONE of the substances, **A** or **B**, is a metal? Give a reason for the answer. (2)

[14]

QUESTION 3 (Start on a new page.)

3.1 The three isotopes of magnesium are Mg-24, Mg-25 and Mg-26. The percentage abundance of the three isotopes is 80%, 10% and 10% respectively.

3.1.1 Define the term *isotope*. (2)

3.1.2 Calculate the relative atomic mass of magnesium. (4)

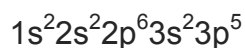
3.1.3 The number of protons and electrons, the mass number and the atomic number of Mg-24 and its ion are shown in the table below.

Some of these values in the table have been omitted. Write down the letters (a–e) in the ANSWER BOOK and next to each letter the number omitted.

	NUMBER OF PROTONS	NUMBER OF ELECTRONS	MASS NUMBER	ATOMIC NUMBER
Mg	12	(a)	24	(b)
Mg ²⁺	(c)	(d)	(e)	12

(5)

3.2 The sp notation of an unknown element X is shown below.



For element X, write down the:

3.2.1 Number of valence electrons (1)

3.2.2 Period where this element is found on the periodic table (1)

3.2.3 Highest energy level in which electrons occur (1)

3.2.4 Symbol (1)

Magnesium combines with element X to form a compound.

3.2.5 Write down the type of bond that forms between magnesium and element X. (1)

3.2.6 Draw the Aufbau diagram for the MAGNESIUM ION. (2)

3.2.7 Draw Lewis dot diagrams to show the bond formation between magnesium and element X. (4)

[22]

QUESTION 4 (Start on a new page.)

A section of the periodic table is shown below. Imaginary symbols are used to represent some of the elements.

	1	2		13	14	15	16	17	18
1									L
2		E					G	J	
3				D		M			
4	A				H				

4.1 Write down the IMAGINARY SYMBOL, as shown above, of the element that:

- 4.1.1 Is a halogen (1)
- 4.1.2 Will form a cation with a charge of +2 (1)
- 4.1.3 Has the largest atomic radius (1)
- 4.1.4 Has the highest electronegativity (1)
- 4.1.5 Is a metalloid (1)
- 4.1.6 Is a noble gas (1)
- 4.1.7 Will form a diatomic molecule (1)
- 4.1.8 Has three valence electrons (1)

4.2 The first ionisation energy of element **A** is $400 \text{ kJ}\cdot\text{mol}^{-1}$.

4.2.1 Define the term *first ionisation energy*. (2)

4.2.2 The first ionisation energy of element **A** can be represented by the following incomplete equation:



Copy the equation above into the ANSWER BOOK and complete it. (2)

4.3 Atoms of element J release the most energy when gaining electrons to form negative ions.

Write down ONE word or term for the underlined phrase. (1)

4.4 Write down the formula of the compound formed when:

4.4.1 **D** combines with **G** (2)

4.4.2 **A** combines with **J** (2)

[17]



QUESTION 5 (Start on a new page.)

Consider the compound $Al_2(SO_4)_3$.

- 5.1 Write down the NAME of the compound above. (1)
- 5.2 Define the term *molar mass*. (2)
- 5.3 Calculate the following for $Al_2(SO_4)_3$:
- 5.3.1 Its molar mass (2)
- 5.3.2 Its percentage composition (3)
- 5.3.3 The number of moles present in 85,5 g (3)
- 5.3.4 The number of aluminium atoms present in 85,5 g (3)
- 5.4 Particles of $Al_2(SO_4)_3$ are bonded in a crystal lattice.
Write down the following for this crystal lattice:
- 5.4.1 The name (1)
- 5.4.2 The particles of which it consists (2)
- 5.4.3 TWO physical properties (2)
- 5.5 500 g of $Al_2(SO_4)_3$ is dissolved in 2 dm³ water.
- 5.5.1 Define the term *concentration*. (2)
- 5.5.2 Calculate the concentration of the solution. (4)

[25]

QUESTION 6 (Start on a new page.)

- 6.1 The first step in the extraction of zinc from zinc sulphide (ZnS) is the combustion of ZnS in oxygen. The balanced equation for the reaction is:



- 6.1.1 Is the reaction above EXOTHERMIC or ENDOTHERMIC? Give a reason for the answer. (2)

Oxygen gas consists of oxygen molecules.

- 6.1.2 Define the term *molecule*. (2)

- 6.1.3 Draw the Lewis dot diagram of the oxygen molecule (O₂). (2)

During the reaction above, 7 g ZnS reacts completely with oxygen gas.

Calculate the:

- 6.1.4 Number of moles of ZnS that has reacted (3)

- 6.1.5 Mass O₂ needed (3)

- 6.1.6 Volume of SO₂(g) produced at STP (4)

- 6.2 Consider the incomplete equation for the chemical reaction below.



- 6.2.1 Write down the NAME of the acid in the equation above. (1)

- 6.2.2 Write down the FORMULA of product X. (2)

- 6.2.3 Name the type of reaction illustrated above. Choose from ACID-BASE REACTION, PRECIPITATION REACTION or REDOX REACTION. Explain the answer. (3)

- 6.2.4 Briefly describe a test that can be used to verify the formation of hydrogen gas during the reaction above. (2)

[24]



QUESTION 7 (Start on a new page.)

7.1 The relationship between conductivity and the concentration of ions in two electrolytes, $\text{NaCl}(\text{aq})$ and $\text{CaCl}_2(\text{aq})$, of the SAME concentration are investigated.

7.1.1 Define the term *electrolyte*. (2)

7.1.2 Is the water molecule POLAR or NON-POLAR? Give a reason for the answer. (2)

7.1.3 For this investigation, write down the:

(a) Independent variable (1)

(b) Dependant variable (1)

The $\text{NaCl}(\text{aq})$ is added dropwise to distilled water in a beaker and the conductivity of the solution is measured after the addition of each drop. The experiment is repeated for the $\text{CaCl}_2(\text{aq})$. The results obtained are shown in the table below.

NUMBER OF DROPS	0	1	2	3	4	5	6	7
Conductivity of $\text{NaCl}(\text{aq})$ (mA)	0,18	0,34	0,55	0,74	0,92	1,10	1,29	1,47
Conductivity of $\text{CaCl}_2(\text{aq})$ (mA)	0,18	0,55	0,91	1,29	1,47	1,84	2,21	2,21

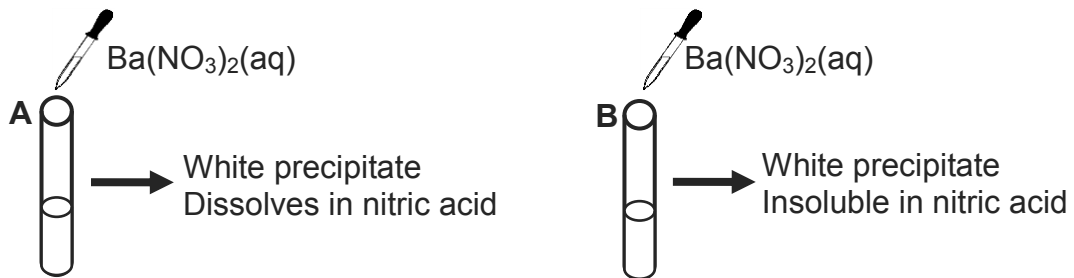
7.1.4 Write down a balanced equation for the dissociation of $\text{NaCl}(\text{s})$ in water. (3)

7.1.5 Which electrolyte, $\text{NaCl}(\text{aq})$ or $\text{CaCl}_2(\text{aq})$, has the higher conductivity? Give a reason for the answer. (2)



- 7.2 A learner is supplied with two unlabelled bottles containing potassium salts. She knows that one bottle contains a SULPHATE and the other a CARBONATE.

To distinguish between the two salts she adds a few drops of barium nitrate, $\text{Ba}(\text{NO}_3)_2(\text{aq})$, to a solution of each salt in two separate test tubes, **A** and **B**, as shown below.



The learner finds that a white precipitate forms in each test tube. After the addition of nitric acid, the precipitate in test tube **A** dissolves to release a gas, while the precipitate in test tube **B** remains.

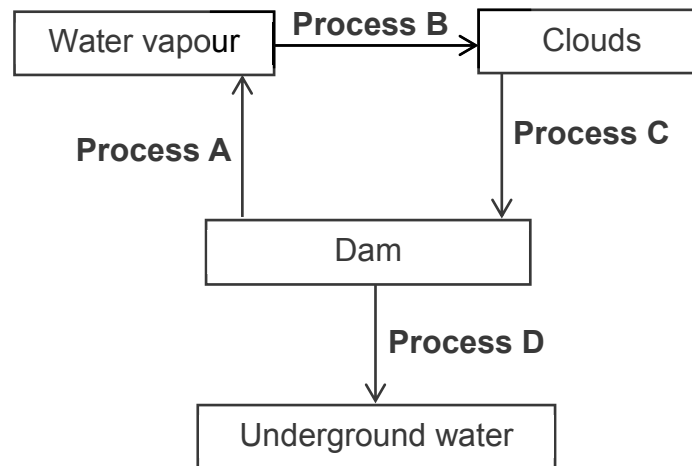
Write down the:

- 7.2.1 FORMULA of the POTASSIUM SALT in test tube **A** (2)
- 7.2.2 FORMULA of the precipitate that forms in test tube **B** (2)
- 7.2.3 Balanced equation for the reaction that explains why the precipitate dissolves in test tube **A** after the addition of nitric acid (3)
- [18]**



QUESTION 8 (Start on a new page.)

The simplified diagram below represents the water cycle. Condensation and evaporation are two of the processes in the water cycle.



- 8.1 Define the term:
- 8.1.1 Condensation (2)
- 8.1.2 Evaporation (2)
- 8.2 Write down the name of:
- 8.2.1 Process A (1)
- 8.2.2 Process B (1)
- 8.2.3 Process C (1)
- 8.2.4 Process D (1)
- 8.3 Describe the term *hydrosphere* with reference to the different phases present. (2)

[10]**TOTAL: 150**

**DATA FOR PHYSICAL SCIENCES GRADE 10
PAPER 2 (CHEMISTRY)**

**GEGEWENS VIR FISIESTE WETENSKAPPE GRAAD 10
VRAESTEL 2 (CHEMIE)**

TABLE 1: PHYSICAL CONSTANTS/TABEL 1: FISIESTE KONSTANTES

NAME/NAAM	SYMBOL/SIMBOOL	VALUE/WAARDE
Standard pressure <i>Standaarddruk</i>	p^θ	$1,013 \times 10^5 \text{ Pa}$
Molar gas volume at STP <i>Molêre gasvolume by STD</i>	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature <i>Standaardtemperatuur</i>	T^θ	273 K
Charge on electron <i>Lading op elektron</i>	e	$1,6 \times 10^{-19} \text{ C}$
Avogadro's constant <i>Avogadro-konstante</i>	N_A	$6,02 \times 10^{23} \text{ mol}^{-1}$

TABLE 2: FORMULAE/TABEL 2: FORMULES

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$
$c = \frac{n}{V}$ OR/OF $c = \frac{m}{MV}$	$n = \frac{V}{V_m}$





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**NATIONAL
SENIOR CERTIFICATE/
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SENIOR SERTIFIKAAT**

GRADE/GRAAD 10

**PHYSICAL SCIENCES: CHEMISTRY (P2)
FISIESE WETENSKAPPE: CHEMIE (V2)**

NOVEMBER 2019

MARKING GUIDELINES/NASIENRIGLYNE

MARKS/PUNTE: 150

**These marking guidelines consist of 10 pages.
*Hierdie nasienriglyne bestaan uit 10 bladsye.***

QUESTION 1/VRAAG 1

- | | | |
|------|------|-------------|
| 1.1 | A ✓✓ | (2) |
| 1.2 | C ✓✓ | (2) |
| 1.3 | B ✓✓ | (2) |
| 1.4 | D ✓✓ | (2) |
| 1.5 | C ✓✓ | (2) |
| 1.6 | C ✓✓ | (2) |
| 1.7 | B ✓✓ | (2) |
| 1.8 | D ✓✓ | (2) |
| 1.9 | C ✓✓ | (2) |
| 1.10 | C ✓✓ | (2) |
| | | [20] |

QUESTION 2/VRAAG 2

2.1

2.1.1 Positive ions/*positiewe ione* ✓
Delocalised valence electrons/*gedelokaliseerde valenselektrone* ✓ (2)

2.1.2 Metallic bond/*metaalbinding* ✓ (1)

2.2 Left/*Links* ✓ (1)

2.3 Hg ✓ (1)

2.4

2.4.1 Homogeneous/*Homogeen* ✓
 Uniform composition/Separate particles cannot be distinguished./All components are in the same phase. ✓
Uniforme samestelling/Afsonderlike deeltjies kan nie onderskei word nie./Alle komponente is in dieselfde fase. (2)

2.4.2 Fe_2O_3 ✓ (1)

2.5

2.5.1 A ✓
Lowest density/*Laagste digtheid* ✓ (2)

2.5.2 Electrical conductivity is the conduction of electric current/charge ✓ and thermal conductivity is the conduction of heat. ✓
Elektriese geleiding is die geleiding van elektriese stroom/ladings en termiese geleiding is die geleiding van hitte. (2)

2.5.3 B ✓
 B has a high density./B is a good conductor of electricity./B is a good conductor of heat. ✓
B het 'n hoë digtheid./B is 'n goeie geleier van elektrisiteit./B is 'n goeie geleier van hitte. (2)

[14]

QUESTION 3/VRAAG 3

3.1

3.1.1

Marking guidelines/Nasiemriglyne

If any of the underlined key words/phrases are omitted: minus 1 mark

Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is: minus 1 punt

Atoms of the same element having the same number of protons, but different number of neutrons. ✓✓

Atome van dieselfde element wat dieselfde getal protone het, maar verskillende getalle neutrone.

OR/OF

Same atomic number, but different mass numbers.

Dieselfde atoomgetalle, maar verskillende massagetalle.

(2)

3.1.2

$$\text{Average/gemiddelde } A_R = \frac{(80)(24)}{100} + \frac{(10)(25)}{100} + \frac{(10)(26)}{100}$$

$$= 24,3$$

(4)

3.1.3

(a) 12

(1)

(b) 12

(1)

(c) 12

(1)

(d) 10

(1)

(e) 24

(1)

3.2

3.2.1 7 ✓

(1)

3.2.2 3 ✓

(1)

3.2.3 3 ✓

(1)

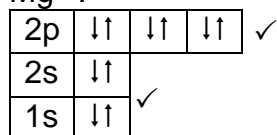
3.2.4 Cl ✓

(1)

3.2.5 Ionic bond/Ioniese binding ✓

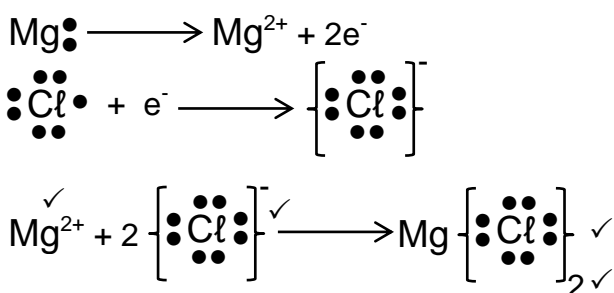
(1)

3.2.6 Mg²⁺:



(2)

3.2.7



(4)

[22]

QUESTION 4/VRAAG 4

- 4.1
- 4.1.1 J ✓ (accept/aanvaar F) (1)
- 4.1.2 E ✓ (accept/aanvaar Be) (1)
- 4.1.3 A ✓ (accept/aanvaar K) (1)
- 4.1.4 J ✓ (accept/aanvaar F) (1)
- 4.1.5 H ✓ (accept/aanvaar Ge) (1)
- 4.1.6 L ✓ (accept/aanvaar He) (1)
- 4.1.7 **Any ONE/Enige EEN**
- G (accept/aanvaar O) ✓
 - J (accept/aanvaar F) (1)
- 4.1.8 D ✓ (accept/aanvaar Al) (1)
- 4.2
- 4.2.1 **Marking guidelines/Nasienriglyne**
If any of the underlined key words/phrases are omitted: minus 1 mark
Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is: minus 1 punt
- First ionisation energy is the energy needed per mole ✓ to remove the first electron from an atom in the gaseous phase. ✓
Eerste ionisasie-energie is die energie benodig per mol om die eerste electron te verwyder vanaf 'n atom in die gasfase. (2)
- 4.2.2 $A(g) + 400 \text{ kJ}\cdot\text{mol}^{-1} \rightarrow \underline{A^+(g)} \checkmark + \underline{e^-} \checkmark$ (2)
- 4.3 Electron affinity/*Elektronaffiniteit* ✓ (1)
- 4.4
- 4.4.1 $D_2G_3 \checkmark\checkmark$ (accept/aanvaar Al_2O_3) (2)
- 4.4.2 AJ ✓✓ (accept/aanvaar KF) (2)
- [17]

QUESTION 5/VRAAG 5

5.1 Aluminium sulphate/Aluminiumsulfaat ✓ (1)

5.2 **Marking guidelines/Nasienriglyne**
If any of the underlined key words/phrases are omitted: minus 1 mark
Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is: minus 1 punt

The mass of one mole of a substance measured in $\text{g}\cdot\text{mol}^{-1}$. ✓✓
Die massa van een mol van 'n stof gemeet in $\text{g}\cdot\text{mol}^{-1}$. (2)

5.3

5.3.1 $M(\text{Al}_2(\text{SO}_4)_3) = 2(27) + 3(32) + 12(16)$
 $= 342 \text{ g}\cdot\text{mol}^{-1}$ ✓✓

Note/Let wel:
If unit omitted/*Indien eenheid uitgelaat*
is: Max./Maks. $\frac{1}{2}$ (2)

5.3.2 **POSITIVE MARKING FROM QUESTION 5.3.1.**
POSITIEWE NASIEN VANAF VRAAG 5.3.1.

$$\% \text{Al} = \frac{2(27)}{342} \times 100$$
$$= 15,79\% \checkmark$$

$$\% \text{S} = \frac{3(32)}{342} \times 100$$
$$= 28,07\% \checkmark$$

$$\% \text{O} = \frac{192}{342} \times 100$$
$$= 56,14\% \checkmark$$
 (3)

5.3.3 **POSITIVE MARKING FROM QUESTION 5.3.1.**
POSITIEWE NASIEN VANAF VRAAG 5.3.1.

$$n(\text{Al}_2(\text{SO}_4)_3) = \frac{m}{M} \checkmark$$
$$= \frac{85,5}{342} \checkmark$$
$$= 0,25 \text{ mol} \checkmark$$
 (3)

5.3.4 **POSITIVE MARKING FROM QUESTION 5.3.3.**
POSITIEWE NASIEN VANAF VRAAG 5.3.3.

Number of Al atoms = $n \times N_A \times$ number of atoms
 $= (0,25)(6,02 \times 10^{23}) \checkmark (2) \checkmark$
 $= 3,01 \times 10^{23}$ atoms ✓ (3)

5.4

5.4.1 Ionic structure/*Ioniese struktuur* ✓ (1)

5.4.2 Al^{3+} / aluminium ions/ positive ions/ *aluminium-ione/positiewe ione* ✓
 SO_4^{2-} / sulphate ions/ negative ions/ *sulfaat-ione/negatiewe ione* ✓ (2)

5.4.3 **ANY TWO/ENIGE TWEE:**

- Brittle/bros ✓
- Hard ✓
- Non-conductor of electricity/ *nie-geleier van elektrisiteit*
- Non-conductor of heat/ *nie-geleier van hitte*
- High melting point/ *hoë smeltpunt*

(2)

5.5

5.5.1 **Marking guidelines/Nasienriglyne**

If any of the underlined key words/phrases are omitted: minus 1 mark
Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is: minus 1 punt

The number of moles of solute per cubic decimetre/litre of solution.
Die getal mol opgelostee stof per kubieke sentimeter/liter van die oplossing.

(2)

5.5.2

<u>OPTION 1/OPSIE 1</u>	<u>OPTION 2/OPSIE 2</u>
$n = \frac{m}{M}$ $= \frac{500}{342} \checkmark$ $= 1,46 \text{ mol}$ $c = \frac{n}{V} \checkmark$ $= \frac{1,46}{2} \checkmark$ $= 0,73 \text{ mol} \cdot \text{dm}^{-3} \checkmark$	$c = \frac{m}{MV} \checkmark$ $= \frac{500}{\sqrt{342}(2)} \checkmark$ $= 0,73 \text{ mol} \cdot \text{dm}^{-3} \checkmark$

(4)

[25]

QUESTION 6/VRAAG 6

6.1

6.1.1 Exothermic ✓
 Energy is released. ✓

Eksotermies.
 Energie word vrygestel.

(2)

6.1.2

Marking guidelines/Nasienriglyne

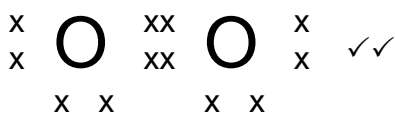
If any of the underlined key words/phrases are omitted: minus 1 mark
Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is: minus 1 punt

A group of two or more atoms that are covalently bonded and that functions as a unit. ✓✓

'n Groep van twee of meer atome wat kovalent gebind is en as 'n eenheid funksioneer.

(2)

6.1.3



(2)

6.1.4
$$n(\text{ZnS}) = \frac{m}{M} \checkmark$$

$$= \frac{7}{97} \checkmark$$

$$= 0,072 \text{ mol} \checkmark$$
 (3)

6.1.5 **POSITIVE MARKING FROM QUESTION 6.1.4.**
POSITIEWE NASIEN VANAF VRAAG 6.1.4.

$$n(\text{O}_2) = \frac{3}{2} n(\text{ZnS})$$

$$= \frac{3}{2} (0,072) \checkmark$$

$$= 0,108 \text{ mol}$$

$$n(\text{O}_2) = \frac{m}{M}$$

$$0,108 = \frac{m}{32} \checkmark$$

$$\therefore m = 3,46 \text{ g} \checkmark$$

Marking criteria/Nasienriglyne

- Use ratio/Gebruik verhouding:

$$n(\text{O}_2) = \frac{3}{2} n(\text{ZnS})$$
- Substitute/Vervang 32 g·mol⁻¹.
- Final answer/Finale antwoord:
3,46 g

6.1.6 **POSITIVE MARKING FROM QUESTION 6.1.4.**
POSITIEWE NASIEN VANAF VRAAG 6.1.4.

$$n(\text{SO}_2) = n(\text{ZnS})$$

$$= 0,072 \text{ mol} \checkmark$$

$$n = \frac{V}{V_m} \checkmark$$

$$0,072 = \frac{V}{22,4} \checkmark$$

$$V(\text{SO}_2) = 1,62 \text{ dm}^3 \checkmark$$

6.2

6.2.1 Sulphuric acid/ swawelsuur ✓

ACCEPT/AANVAAR:

Hydrogen sulphate/Waterstofsulfaat ✓

6.2.2 ZnSO₄ ✓✓

6.2.3 Redox (reaction)/Redoks(reaksie) ✓

The charge of Zn changes from 0 in Zn ✓ to +2 in ZnSO₄. ✓ /The charge of H changes from +1 in H₂SO₄ to 0 in H₂.

Die lading van Zn verander van 0 in Zn na +2 in ZnSO₄. /Die lading van H verander van + 1 in H₂SO₄ na 0 in H₂.

6.2.4 When a burning woodsplinter is brought close to the gas it makes a popping sound. ✓✓

'n Brandende houtsplinter wat naby die gas gebring word, maak 'n plofgeluid.

(2) [24]

QUESTION 7/VRAAG 7

7.1

7.1.1

Marking guidelines/Nasienriglyne

If any of the underlined key words/phrases are omitted: minus 1 mark

Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is: minus 1 punt

An electrolyte is a solution that conducts electricity ✓ through the movement of ions. ✓

'n Elektroliet is 'n oplossing wat elektrisiteit gelei deur die beweging van ione. (2)

7.1.2 Polar/Polêr ✓

It has two oppositely charged poles/Dit het twee teenoorgesteld gelaaide pole. ✓ (2)

7.1.3

a) Concentration of ions/Konsentrasie van ione ✓ (1)

b) Conductivity/Geleidingsvermoë ✓ (1)

7.1.4 $\text{NaCl(s)} \xrightarrow{\text{H}_2\text{O(l)}} \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ ✓ Bal. ✓

NOTE/LET WEL:

Ignore phases./Ignoreer fases. (3)

7.1.5 CaCl_2 ✓



A higher concentration of ions forms in solution. ✓
'n Hoër konsentrasie van ione vorm in oplossing. (2)

7.2

7.2.1 K_2CO_3 ✓✓ (2)

7.2.2 BaSO_4 ✓✓ (2)

7.2.3 $\text{BaCO}_3(\text{s}) + 2\text{HNO}_3(\text{aq}) \rightarrow \text{Ba}(\text{NO}_3)_2(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ ✓ Bal. ✓

NOTE/LET WEL:

Ignore phases./Ignoreer fases. (3)

[18]

QUESTION 8/VRAAG 8

8.1

8.1.1

Marking guidelines/Nasienriglyne

If any of the underlined key words/phrases are omitted: minus 1 mark
*Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is:
minus 1 punt*

Condensation is the process during which a gas or vapour changes to a liquid. ✓✓

Kondensasie is die proses waardeur 'n gas of damp in 'n vloeistof verander.

(2)

8.1.2

Marking guidelines/Nasienriglyne

If any of the underlined key words/phrases are omitted: minus 1 mark
*Indien enige van die onderstreepte sleutelwoorde/frases uitgelaat is:
minus 1 punt*

Evaporation is the change of a liquid into a vapour at any temperature below the boiling point. ✓✓

Verdamping is die verandering van 'n vloeistof in 'n damp by enige temperatuur onder die kookpunt.

(2)

8.2

8.2.1 Evaporation/Verdamping ✓

(1)

8.2.2 Condensation/Kondensasie ✓

(1)

8.2.3 Precipitation/Presipitasie ✓

(1)

8.2.4 Infiltration/Infiltrasie/insypeling ✓

(1)

8.3 The hydrosphere is the water of the Earth. ✓ It is found as liquid water, ice and water vapour in the atmosphere. ✓

Die hidrosfeer is die water van die Aarde en dit word gevind as vloeibare water, ys en waterdamp in die atmosfeer.

(2)

[10]

TOTAL/TOTAAL:

150

